Study guide 8P1.1

1. 11.24 g/ml
2. 1.33 g/cm3
3. 5.2 g/ml
4. .68 L
5. 347.4 g
6. 11.4 g
7. Lx w x h
8. Soap – M

Smoke – m

Heat – n

Hate – n

Atoms – m

Books – m

Air – m

Light – n

You – m

Ice – m

Sand – m

Sound – n

1. Does it have mass? Does it take up space?
2. Proton – nucleus

Neutron – 0 no charge

* Electron cloud – out side of nucleus

1. Strawberry, atom, nucleus, proton, neutron, electron
2. A – electron

B – neutron

C – proton

D – nucleus

13 a. atom

13.b matter

13.c heterogeneous

13d. volume

13e. element

13f. compound

13g. mass

13h. mixture

14.a mixture of compounds

14b. element

14c. pure compound

14d. compound

14e. mixture of elements

14f. mixture of compounds and elements

15.a A – packed close together in a pattern. Has a shape

15b. B – more fluid – takes on shape of container

15c. C – atoms spreading apart and filling the container

15d. A has the highest density because it is a solid

15e. density decreases

1. Sand – mixture = heterogeneous
2. Salt – compound
3. Pure Water – compound
4. Soil – mixture – heterogeneous
5. Soda just opened – mixture – homogeneous
6. Pure air – compound
7. Carbon dioxide = compound
8. Gold – element
9. Brass - Compound